

Sample				
9	9	1	2	3
0	0	0	0	0

0	0	0	0	0
1	1	1	1	1
2	2	2	2	2
3	3	3	3	3
4	4	4	4	4
5	5	5	5	5
6	6	6	6	6
7	7	7	7	7
8	8	8	8	8
9	9	9	9	9

33 P-597

SSC-A-I-2026 (9th)

0+0=

Roll No
Student Must Fill the
Roll Number Bubbles.

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991234



(Section-A) Marks= (12) Time Allowed: 20 Minutes

All questions of this section must be answered on this page and handed over to the Superintendent within given time

Student Name

Name of Paper CHEMISTRY (CH-I)

Paper Date 13-APR-26

165

Paper Code: 73

Signature of Student: _____

Paper Version

A B C

Sign and Seal of Suptd: _____ SEAL



0	0	0	0	0
1	1	1	1	1
2	2	2	2	2
3	3	3	3	3
4	4	4	4	4
5	5	5	5	5
6	6	6	6	6
7	7	7	7	7
8	8	8	8	8
9	9	9	9	9

سوالات

سوال نمبر کے مطابق درست جواب نمبر کا دائرہ پر کریں۔

- Vapor pressure of a liquid depends mainly on: a. amount of liquid, b. shape of container, c. nature of liquid and temperature, d. colour of liquid
- The molality of a solution containing 2.0 moles of solute in 1.0 dm³ of solution is: a. 0.5 M, b. 1.0 M, c. 2.0 M, d. 4.0 M
- Milk is an example of: a. solution, b. colloid, c. suspension, d. solid solution
- The oxidation number of sodium (Na) in NaCl is: a. -1, b. 0, c. +1, d. +2
- The particles which move towards the anode in an electrolytic cell are: a. cations, b. anions, c. protons, d. neutrons
- The first ionization energy of alkaline earth metals is higher than alkali metals because alkaline earth metals have: a. larger atomic size, b. more shielding effect, c. higher nuclear charge, d. fewer protons
- The preparation of cement, glass and fertilizer is studied in: a. industrial chemistry, b. organic chemistry, c. biochemistry, d. nuclear chemistry
- The following compound has the same empirical and molecular formula: a. benzene, b. glucose, c. acetic acid, d. water
- Protium (¹H) contains: a. 0 neutron, b. 1 neutron, c. 2 neutrons, d. 3 neutrons
- Across a period from left to right, electronegativity generally: a. decreases, b. remains same, c. increases, d. becomes zero
- The element which follows the duplet rule for stability is: a. neon, b. argon, c. oxygen, d. helium
- Ionic compounds are usually soluble in: a. non-polar solvents, b. polar solvents like water, c. oils, d. benzene

- (A) (B) (C) (D)
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If the Candidate is absent, The Superintendent Should fill this Circle and pack the sheet with solved OMR sheets in roll number order

CHEMISTRY — 9th

SSC-9th-Abbottabad-2026

Time: 2 Hours 40 Minutes

SECTION-B**Marks: 32****NOTE:** Attempt any eight of the following short questions. Each question carries 4 marks

- Q.1. Classify the following as element, compound or mixture by giving reason in each case.
(i) Iron (ii) Carbon dioxide (iii) Salt solution (iv) Oxygen 1+1+1+1
- Q.2. Write down the electronic configuration of the following elements:
(i) $_{11}\text{Na}$ (ii) $_{13}\text{Al}$ (iii) $_{16}\text{S}$ (iv) $_{17}\text{Cl}$ 1+1+1+1
- Q.3. How does shielding effect affect the ionization and electron affinity of an element? 2+2
- Q.4. Explain a triple covalent bond. How is it formed in nitrogen (N_2) molecule. 2+2
- Q.5. Describe the process of anion formation in non-metals. Explain how does an oxygen atom form an oxide ion (O^{2-}) 2+2
- Q.6. Draw and explain the graph obtained by plotting volume versus absolute temperature for a gas at constant pressure. 2+2
- Q.7. What is supersaturated solution? How is it prepared? 2+2
- Q.8. Explain why is sodium chloride (NaCl) soluble in water but insoluble in benzene, using the rule "like dissolves like". 2+2
- Q.9. Explain the role of alloying in corrosion prevention. Give one example of a corrosion resistant alloy. 2+2
- Q.10. State the half reactions at anode and cathode during electroplating of tin on steel. 2+2
- Q.11. Compare the commercial importance of silver and gold by stating two uses of each metal. 1+1+1+1

SECTION-C**Marks: 21****NOTE:** Attempt any three of the following questions. Each question carries 7 marks

- Q.12. i. Describe four main points of Rutherford's atomic model. 1+1+1+1
ii. Calculate the number of atoms present in 3 moles of aluminium (Al). 1+1+1
- Q.13. i. Differentiate between ionic and covalent bonds with one example each. 2+2
ii. What is periodic table? Explain with example how do the number of electron shells and valence electrons determine the position of an element in periodic table. 1+1+1
- Q.14. i. What is liquid solution? Explain the formation of liquid in liquid solution with example. 2+2
ii. Define melting point. Why do crystalline solids have sharp melting point and amorphous solids don't. 1+1+1
- Q.15. i. Draw a diagram of a Daniel cell and label anode and cathode. Also write from which electrode electrons move to which electrode. 1+2+1
ii. Write the balanced equation when chlorine reacts with hydrogen in the presence of sunlight. Name the product formed and justify the need of sunlight for this reaction. 1+1+1